

## Research Article

# Breakdown of PFOA, PFOS and 6:2FTS Using Acidic Potassium Permanganate as Oxidant

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**Abstract**

We found that inert fluoro-carbon skeletons of Perfluorooctanoic Acid (PFOA), Perfluorooctane Sulfonate (PFOS) and 1H, 1H, 2H, 2H-perfluorooctanesulfonic acid (6:2FTS) could be broken down by potassium permanganate as oxidant in acidic liquid phase at room temperature. This opened a new Approach to the Remediation of Aqueous Film Forming Foams (AFFFs). The breakdown was confirmed from the HPLC-MS and Ion Chromatography (IC) data. Due to the oxidization's contribution, those fluoro-carbon skeletons' half-life was estimated to be approximately 3 months, much shorter than the several decades that occur in nature.

**Keywords:** PFOA; PFOS; 6:2FTS; Breakdown; Oxidization; Potassium permanganate

**Introduction**

Poly- and Perfluoroalkyl Substances (PFASs) exhibit unique physical and chemical properties, such as hydrophobicity and oleophobicity, which are not evident in other components, and also extreme stability with respect to thermal, chemical and biodegradation [1]. Due to their important anti-wetting and anti-staining properties, they have been used widely and domestically in such activities as clothing, upholstery, carpeting, painted surfaces, food containers, cookware, etc [2]. However, since their fluoro-carbon skeletons are inert and resistant to biodegradation under natural environmental conditions ( $\text{CF}_3\text{-CF}_3$  of 99 kcal/mol vs  $\text{CH}_3\text{-CH}_3$  of 89 kcal/mol) [3,4] this has led to their global accumulation and distribution in the environment. This has in turn raised serious concerns about their impact on the environment and public health [5-7].

Aqueous Film Forming Foam (AFFF) is a good example that has been widely used to extinguish fires [8,9]. Its main ingredients are anionic fluoro surfactants such as Perfluorooctane Sulfonate (PFOS) and perfluorooctanoic acid (PFOA). Due to serious misgivings about their biological and environmental impact and their persistent nature, PFOS was phased out in the early 2000s. Lots of alternatives were found on the market. These include, for example, 1H, 1H, 2H, 2H-perfluorooctanesulfonic acid (6:2FTS)-, and 1H, 1H, 2H, 2H-perfluorodecane sulfonic acid (8:2FTS)-based fluoro surfactants [10]. Those fluorotelomers were synthesised *via* telomerisation with linear structures that differ from the products of electrochemical fluorination, such as PFOS containing linear and branched isomers [6,11]. Although they are re-reported to be environmentally safe their fluoro-carbon skeletons still raise concerns about their biodegradability in the natural environment [12]. For example, the half-life of 6:2FTS is estimated to be >10 years [12], which is shorter but still similar to >41 years for PFOS, and >92 years for PFOA (USEPA 505-F-14-001), respectively. It should be noted that those values depend on estimating approach, initial concentration, degradation conditions, etc., and consequently they have varied in the literature [12-14]. New ingredients thus include short chains of the fluoro-

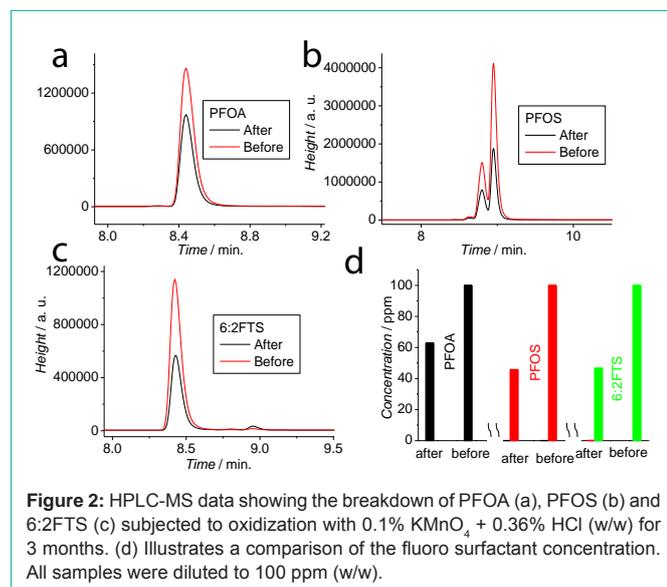
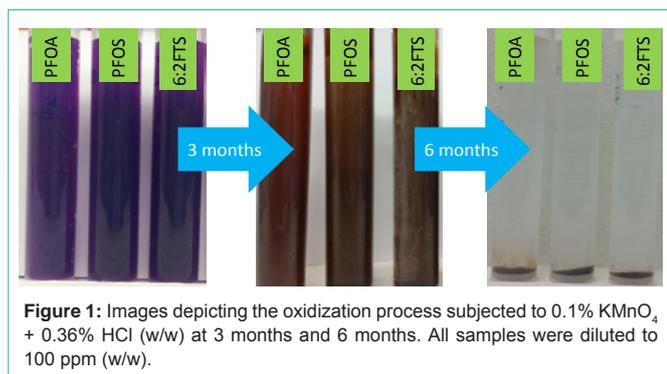
carbon ( $\text{C}_3\text{-C}_6$ ) [15,16] and fluorine-free surfactants [17].

Previously we used chemical oxidant (potassium permanganate,  $\text{KMnO}_4$ ) to break down the derived groups of fluoro surfactants because the non-fluoro-carbon can be broken down much more easily than fluoro-carbon skeleton [18]. However, we found that the fluoro-carbon skeleton could potentially be broken down, although the process was slow. This particular phenomenon is interesting because it might lead to a new degradation approach that is different from previous ones, for example Fenton reaction of hydrogen peroxide [19], persulphate [4], advanced electrochemical oxidization [20], Sonolytic conversion [21] etc. Compared to the heat-up approach [22], this mild condition (occurred at room temperature) offers the promise to scale-up its application. The additional advantages of this oxidant include cost-effectiveness, stability, environmentally safe, easy to operate etc [23,24]. Here we selected 3 common fluoro surfactants - PFOA, PFOS and 6:2FTS - to verify the possibility of breaking down these ingredients using  $\text{KMnO}_4$ .

**Materials and Methods**

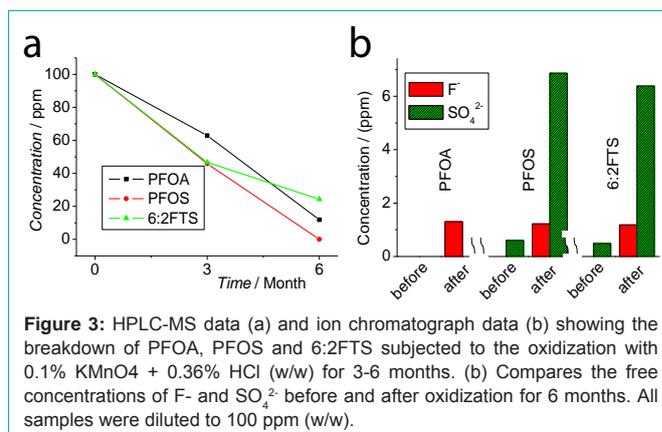
All chemicals including PFOA, PFOS and 6:2FTS, potassium permanganate ( $\text{KMnO}_4$ , ACS reagent,  $\geq 99.0\%$ ), hydrogen chloride (HCl, 37%, w/w, AR), methanol and ammonium acetate ( $\text{NH}_4\text{Ac}$ ) were purchased from Sigma-Aldrich (Australia). Only polypropylene containers/pipette tips were used throughout to avoid any potential interference from Teflon containers/caps. Milli-Q water was used ( $> 18 \text{ M}\Omega\cdot\text{cm}$ ) in the present study.

All samples were diluted in Milli-Q water in centrifuge tubes (polypropylene) without pre-treatment. 0.1%  $\text{KMnO}_4$  + 0.36% HCl (w/w) was placed in the tubes for the oxidization process [4,23,24]. The tubes were kept at room temperature ( $\sim 24^\circ\text{C}$ ) and not shielded from the laboratory fluorescent lamp for the purposes of domestic lighting. The tubes were occasionally shaken (once per day) during oxidization. Samples were filtered with nylon syringe filters (0.2  $\mu\text{m}$ ) prior to HPLC-MS analysis HPLC-MS (Agilent 1260 + Quadrupole 6130) before and after the oxidization [25,26].



For HPLC-MS analysis, we followed the standard method (EPA/600/R-08/092) [27]. In general, 10  $\mu\text{L}$  sample solution was injected into Agilent 1260 high-performance liquid chromatography fitted with an XDB- $\text{C}_{18}$  column kept at 40 °C. Its dimensions were 2.1 mm internal diameter, 100 mm length and 5  $\mu\text{m}$  particle size. The flow rate was 0.5 mL/min for gradient mobile phase of methanol: 5 mM aqueous  $\text{NH}_4\text{Ac}$  for separation. Quadrupole 6130 detector was maintained at 70 V under negative mode for scanning. Extraction of the molecular ions was conducted at  $m/z$  413 for PFOA, 499 for PFOS and 427 for 6:2FTS, respectively. Quantification was done by producing a calibration curve using external standard solutions of PFOA, PFOS (only linear isomers were quantified) and 6:2FTS with correlation coefficients higher than 0.99 and limit of detection being  $\sim 0.2$  ppb (signal: noise > 3). Blank samples of Milli-Q water and methanol were run prior to each set of test to minimize any background contamination that could have originated from the Teflon components of the HPLC instrument itself. The nebulizer gas (nitrogen) pressure was set at 40 psi, drying gas flow rate was 9 L/min and temperature 325 °C, capillary voltage was + 3500 V and skimmer voltage was - 15 V. More details are listed in Ref [21].

Free fluoride ( $\text{F}^-$ ) and sulphate ( $\text{SO}_4^{2-}$ ) were detected using Ion Chromatography (IC), which was conducted using DIONEX (ICS-2000, RFIC, and Thermo Scientific). The ion exchange column was IonPac™ AS18, 2 × 250 mm and kept at 35 °C under 2230 psi



pump pressure. Following 25  $\mu\text{L}$  sample injection, 10 mM  $\text{KOH}$  was gradually flow at 0.25 mL/min. Conductivity detector was employed with a suppressor of 43 mA.

Note that each time we ran at least 6 samples in parallel (3 samples without addition of oxidant as controls and the other 3 samples with oxidant) for quality assurance and quality control (QA/QC) [28].

## Results and Discussion

Figure 1 indicates that the colour change depends on the oxidation process. At the beginning, the purple solution confirms the existence of  $\text{KMnO}_4$ . The colour became increasingly darker and changed to brown after 3 months, suggesting the decomposition of oxidant  $\text{KMnO}_4$ . With this decomposition some targets have been oxidized, such as PFOA, PFOS or 6:2FTS in the solution, although the nature of the decomposition of  $\text{KMnO}_4$  should not be ignored [23,24].

After 6 months, the solution became transparent and some precipitates were observed on the bottom of the containers, these being the products of the oxidation, such as  $\text{MnSO}_4$ ,  $\text{MnO}_2$ , etc. Those precipitates should not be injected into the HPLC column so that they need to be filtered off for HPLC analysis.

Figure 2 shows the HPLC-MS results when fluoro surfactants were subjected to the oxidation using 0.1%  $\text{KMnO}_4$  + 0.36%  $\text{HCl}$  (w/w). All 3 samples, including PFOA (a), PFOS (b) and 6:2FTS (c), feature the decreased peak heights after the oxidation (marked as “after”) compared to the absence of oxidant in the solution as controls (marked as “before”), suggesting the breakdown of the fluoro surfactants. (d) high lights the quantitative data where PFOA, PFOS and 6:2FTS were calibrated using external standards. Here the multi-peaks in (b) are assigned to its isomers. We just analysed the linear isomer using its standard. Basically, after 3 months’ oxidation, a significant decrease in the 3 samples’ concentration, from 62% to 45%, was observed.

Figure 3 shows the HPLC-MS data and IC data after 6 months’ oxidation. In Figure 3 (a), we can see the concentrations of fluorosurfactants declined further compared to the results of 3 months’ oxidation, suggesting that the oxidation process was gradual. The kinetics information was unclear from the limited data points. However, we can estimate the half-life to be approximately 3 months for all 3 samples, which is much shorter than the several

decades that occur in the natural environment [1,29], suggesting that the oxidant does make a contribution. We should also note that the oxidization capacity of 0.1%  $\text{KMnO}_4$  + 0.36% HCl faded with time and was not as strong as the fresh solution, as observed in Figure 1. In other words, the breakdown process can be accelerated if the oxidant solution can be refreshed continuously. Another option is the temperature increase that warrants further research.

Figure 3 (b) shows the IC data for the free ions of  $\text{F}^-$  and  $\text{SO}_4^{2-}$ . We can see the concentration of  $\text{F}^-$  significantly increased after 6 months' oxidization when compared to almost zero (less than the limit of detection, 0.5 ppm) for all 3 samples in the absence of oxidization. It confirmed the breakdown of the fluoro-carbon skeleton of the fluorosurfactants because the free ion of  $\text{F}^-$  was released from the fluoro-carbon skeleton. For PFOS and 6:2FTS, we also observed the increased concentration of  $\text{SO}_4^{2-}$  after oxidization, which originated from their sulfonic groups. Conversely, there is no detectable  $\text{SO}_4^{2-}$  before or after the oxidization of PFOA, supporting the above assumption that sulfonic-containing groups were broken down and converted into  $\text{SO}_4^{2-}$  because PFOA does not contain this kind of group. Note that the  $\text{SO}_4^{2-}$  was detected before oxidization commenced, which may be due to the impurity from the original sample, or due to the partial breakdown by oxygen from air because we kept the control samples (without oxidant) in parallel including dilution and storage in air for 6 months.

The concentrations of  $\text{F}^-$  were estimated to be ~1.2 ppm (~0.6 mM) for all 3 samples and  $\text{SO}_4^{2-}$  6.5 ppm (~0.068 mM) for PFOS and 6:2FTS, respectively. Considering the samples have been diluted to 100 ppm (0.24 mM for PFOA, 0.20mM for PFOS and 0.23mM for 6:2FTS), the measured concentration of  $\text{F}^-$  and  $\text{SO}_4^{2-}$  indicated the breakdown of fluoro-carbon is not 100% but 17-20% in terms of  $\text{F}^-$ , 25-30% in terms of  $\text{SO}_4^{2-}$  (taking out the amount before the oxidization and only for PFOA and 6:2FTS), respectively. This agrees with the results in Figure 3 (a). Another possible reason is the fluoro-carbon skeleton were mainly broken down into fragments, but not completely into free inorganic ions, which requires more experimental research. The leakage of HF gas into air should also be considered.

In summary, we successfully demonstrated the degradation of PFOA, PFOS and 6:2FTS using  $\text{KMnO}_4$  as oxidant at the room temperature. Although the kinetics information is absent, the half-life for all 3 fluorosurfactants was estimated to be ~3 months, which represents a rapid approach compared to natural degradation of 870-1400 years or 10-17 years for surface-mediation [12]. Compared to 30-60 min for advanced electrochemical oxidization and catalyzed  $\text{H}_2\text{O}_2$  propagation reactions [19], this approach is simple and cost-effective for breaking down fluoro surfactant samples. Admittedly, the concentrations of PFOA, PFOS and 6:2FTS in this test are higher than the ones in nature, such as in contaminated groundwater [31]. Therefore, a pre-concentration module might be needed to improve the breakdown efficiency for the practical remediation.

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